

PRACTICE EXERCISE 2.1

Q1. What colour do the following indicators turn when added to a base or alkali (such as sodium hydroxide)

(a) methyl orange (b) litmus
(c) red cabbage extract

Q2. What colours do the following indicators turn when added to an acid (such as hydrochloric acid)?

(a) litmus (b) Methyl orange

Q3. Name an indicator which is red in acid solution but turn blue in basic solution.

Q4. Name an indicator which is pink in alkaline solution but turns colourless in acidic solution.

Q5. When a solution is added to a cloth strip treated with onion extract, then the smell of onion cannot be detected. State whether the given solution contains an acid or a base.

Q6. When a solution is added to vanilla extract, then the characteristic smell of vanilla cannot be detected. State whether the given solution is an acid or a base.

Q7. How will you test for the gas which is liberated when hydrochloric acid reacts with an active metal?

Q8. Name the gas evolved when dilute HCl reacts with sodium hydrogencarbonate. How is it recognised?

Q9. Give the names and formulae of two strong acids and two weak acids.

Q10. Name one natural source of each of the following acids:

(a) Citric acid (b) Oxalic
(c) Lactic acid (d) Tartaric acid

Q11. Name one animal acid one plant whose stings contain formic acid (or methanoic acid).

Q12. How is the concentration of hydronium ions (H_3O^+) affected when the solution when the

solution of an acid is diluted?

Q13. Write word equations and then balanced equations for the reactions taking place when:

(a) dilute sulphuric acid reacts with zinc granules.
(b) dilute hydrochloric acid reacts with magnesium ribbon.
(c) dilute sulphuric acid reacts with aluminium powder.
(d) dilute hydrochloric acid reacts with iron filings.

Q14. Complete and balance the following chemical equations:

(a) $Zn(s) + HCl(aq) \longrightarrow$
(b) $Na_2CO_3(s) + HCl(aq) \longrightarrow$
(c) $NaHCO_3(s) + HCl(aq) \longrightarrow$
(d) $NaHCO_3(s) + HCl(aq) \longrightarrow$
(e) $CuO(s) + HCl(aq) \longrightarrow$

Q15. Fill in the blanks in the following sentences:

(a) Acids have a taste and they turn litmus to
(b) Substances do not show their acidic properties without
(c) Those substances whose smell (or odour) changes in acidic or basic solutions are called indicators.
(e) Onion and vanilla extract are indicators.

Q16. What is an indicator? Name three common indicators.

Q17. Name the acid-base indicator extracted from lichen.

Q18. What colour does the turmeric paper turn when put in an alkaline solution?

PRACTICE EXERCISE 2.2

Q1. What is an olfactory indicator? Name two olfactory indicators. What is the effect of adding sodium hydroxide solution to these olfactory indicators?

Q2. What happens when an acid reacts with a metal? Give chemical equations of the reaction involved.

Q3. Which gas is usually liberated when an acid reacts with a metal? How will you test for the presence of this gas?

Q4. While diluting an acid, why is it recommended that the acid should be added to water and not water to the acid?

Q5. What happens when an acid reacts with a metal hydrogencarbonate? Write equations of the reaction which takes place.

Q6. What happens when dilute hydrochloric acid is added to sodium carbonate? Write a balanced chemical equation of the reaction involved.

Q7. Which gas is liberated when dilute hydrochloric acid reacts with sodium carbonate? How will you test for the presence of this gas?

Q8. What happens when an acid reacts with a base? Explain by taking the example of hydrochloric acid and sodium hydroxide. Give equation of the chemical reaction which takes place. What is the special name of such a reaction?

Q9. What happens when an acid reacts with a metal oxide? Explain with the help of an example. Write a balanced equation for the reaction involved.

Q10. What are organic acids and mineral acids?

Q11. Give two examples each of organic acids and mineral acids.

Q12. State some of the uses of mineral acids in industry.

Q13. What is meant by strong acids and weak acids? Classify the following into strong acids and weak acids:

Q14. Why do HCl , H_2SO_4 , HNO_3 , etc., show acidic character in aqueous solutions while solutions of compounds like $\text{C}_6\text{H}_{12}\text{O}_6$ (glucose) and $\text{C}_2\text{H}_5\text{OH}$ (alcohol) do not show acidic character?

Q15. What is a neutralisation reaction? Explain with an example. Give the chemical equation of the reaction which takes place.

Q16. Why should curd and other sour foodstuffs (like lemon juice, etc.) not be kept in metal containers (such as copper and brass vessels)?

Q17. What is produced if an acid is added to a base?

Q18. Why does dry HCl gas not change the colour of dry litmus paper?

Q19. What colour does phenolphthalein indicator turn when added to an alkali (such as sodium hydroxide)?

Q20. Why do acids not show acidic behaviour in the absence of water?

Q21. Why does an aqueous solution of an acid conduct electricity?

Q22. Why does distilled water not conduct electricity whereas rain water does?

Q23. What happens when an acid reacts with a metal carbonate? Explain with the help of an example. Write chemical equation of the reaction involved.

Q24. What happens when carbon dioxide gas is passed through lime water: (i) for a short time? (ii) for a considerable time? Write equations of the reactions involved.

Q25. With the help of labelled diagrams, diagrams, describe an activity to show that acids produce ions only in aqueous solutions.

Q26. Which element is common to all acids?

PRACTICE EXERCISE 2.3

Q1. When a piece of limestone reacts with dilute HCl, a gas X is produced. When gas X is passed through lime water then a white precipitate Y is formed. On passing excess of gas X, the white precipitate dissolves forming a soluble compound Z.

- What are X, Y and Z?
- Write equations for the reactions which take place:
 - when limestone reacts with dilute HCl
 - when gas X reacts with lime water to form white precipitate Y
 - when excess of gas X dissolves white precipitate Y to form a soluble compound Z.

Q2. If someone is suffering from the problem of acidity after overeating, which of the following would you suggest as remedy?

Lemon juice, Vinegar, Baking soda solution
Give reason for your choice.

Q3. On adding dilute hydrochloric acid to copper oxide powder, the solution formed is blue-green.

- Predict the new compound formed which imparts a blue-green colour to solution.
- Write a balanced chemical equation of the reaction which takes place.
- On the basis of the above reaction, what can you say about the nature of copper oxide?

Q4. A white shirt has a yellow stain of curry. When soap is rubbed on this shirt during washing, the yellow stain turns reddish-brown. On rinsing the shirt with plenty of water, the reddish-brown stain turns yellow again.

- Name the natural indicator present in curry stain.

Q5. You have been provided with three test-tubes. One of these test-tubes contains distilled water and the other two contain an acidic and a basic solution respectively. If you are given only blue litmus paper, how will you identify the contents of each test-tube?

Q6. A substance X which is used as an antacid reacts with dilute hydrochloric acid to produce a gas Y which is used in one type of fire-extinguisher. Name the substance X and gas Y. Write a balanced equation for the chemical reaction which takes place.

Q7. How is the neutralisation of a carbonate with an acid different from the neutralisation of an oxide or a hydroxide?

Q8. What happens to (a) the H^+ ions, and (b) temperature of the solution, when an acid is neutralised?

Q9. What will happen if hydrochloric acid is poured over egg shells.

Q10. How can you test that metal oxides are basic in nature.

Q11. What acids are present in oranges and tomatoes.

Q12. Name two acids used in food stuff.

Q13. How can you test CO_2 is acidic in nature.

Q14. What is the hardest substance in our body and what is it chemically.

Q15. Why dull surface of copper starts shining when we rub it with lemon.

Q16. What is the colour of the litmus solution extracted from the lichen plant.

PRACTICE EXERCISE 2.4

Q1. 10 ml. of a solution of NaOH is found to be completely neutralised by 8 ml. of a given solution of HCl. If we take 20 ml. of the same solution of NaOH, the amount of HCl solution (the same solution as before) required to neutralise it will be:

(a) 4 ml. (b) 8 ml.
(c) 12 ml. (d) 16 ml.

Q2. Which of the following types of medicine is used for treating indigestion caused by overeating?

(a) antibiotic (b) analgesic
(c) antacid (d) antiseptic

Q3. A solution reacts with marble chips to produce a gas which turns lime water milky. The solution contains.

(a) Na_2SO_4 (b) CaSO_4
(c) H_2SO_4 (d) K_2SO_4

Q4. One of the following is not an organic acid. This is:

(a) ethanoic acid (b) formic acid
(c) citric acid (d) carbonic acid

Q5. The property which is not shown by acids is:

(a) they have sour taste
(b) they feel soapy
(c) they turn litmus red
(d) their pH is less than seven

Q6. The indicators which turn red in acid solution are:

(a) turmeric and litmus
(b) phenolphthalein and methyl orange
(c) litmus and methyl orange
(d) phenolphthalein and litmus

Q7. The discomfort caused by indigestion due to

overeating can be cured by taking:

(a) vinegar (b) lemon juice
(c) baking soda (d) caustic soda

Q8. The property which is common between vinegar and curd is that they:

(a) have sweet taste
(b) have bitter taste
(c) are tasteless
(d) have sour taste

Q9. The indicator which produces a pink colour in an alkaline solution is:

(a) methyl orange (b) turmeric paper
(c) phenolphthalein (d) litmus paper

Q10. A solution reacts with zinc granules to give a gas which burns with a 'pop' sound. The solution contains

(a) $\text{Mg}(\text{OH})_2$ (b) Na_2CO_3
(c) NaCl (d) HCl

Q11. Which of the following will turn red litmus to blue:

(a) vinegar (b) baking soda
(c) lemon juice (d) soft drink

Q12. Solutions A, B, C and D have pH 3, 4, 6 and 8. The solution with highest acidic strength is

(a) A (b) B
(c) C (d) D

Q13. Methyl orange is

(a) red in acidic medium, yellow in basic solution
(b) yellow in acidic solution, red in basic solution
(c) Colourless in acidic solution, red in basic solution
(d) Red in acidic solution, colourless in basic solution

PRACTICE EXERCISE 2.5

Q1. Name the gas evolved when zinc granules are treated/heated with:

- hydrochloric acid solution
- sodium hydroxide solution

Q2. What is the common name of water soluble bases?

Q3. What is the common in all the water soluble bases (or alkalis)?

Q4. Why does tooth decay start when the pH of mouth is lower than 5.5?

Q5. What is the pH of a neutral solution?

Q6. Which is more acidic: a solution of pH = 2 or a solution of pH = 6?

Q7. Which is more basic (or more alkaline): a solution of pH = 8 or a solution of pH = 11?

Q8. Name the scientist who developed the pH scale.

Q9. Name the indicator which can give us an idea of how strong or weak an acid or base is.

Q10. The pH of soil A is 7.5 while that of soil B is 4.5. Which of the two soils, A or B, should be treated with powdered chalk to adjust its pH and why?

Q11. What is the name of the indicator which can be used for testing the pH of a solution.

Q12. What colour will universal indicator show if you add it to the following substances?

- potassium hydroxide, pH = 12
- soda water, pH = 5
- sulphuric acid, pH = 2

Q13. A beaker of concentrated hydrochloric acid has a pH of 1. What colour will full range universal indicator turn if it is added to this beaker? Is it a strong or a weak acid?

Q14. Two solutions X and Y are tested with universal indicator. Solution X turns orange whereas solution Y turns red. Which of the solutions is a stronger acid?

Q15. Two solutions A and B have pH values of 3.0 and 9.5 respectively. Which of these will turn litmus solution from blue to red and which will turn phenolphthalein from colourless to pink?

Q16. Two drinks P and Q gave acidic and alkaline reactions, respectively. One has a pH value of 9 and the other has a pH value of 3. Which drink has the pH value of 9?

Q17. Two solutions X and Y have pH = 4 and pH = 8, respectively. Which solution will give alkaline reaction and which one acidic?

Q18. Fill in the following blanks with suitable words:

- Acids have a pH than 7.
- Alkalies have a pH than 7.
- Neutral substances have a pH of
- The more acidic a solution, the the pH.
- The more alkaline a solution, the the pH.

Q19. Fresh milk has a pH of 6. When it changes into curd (yogurt), will its pH value increase or decrease? Why?

Q20. What is a universal indicator? For what purchase is it used?

Q20. How does a universal indicator work?

Q21. Water is a neutral substance. What colour will you get when you add a few drops of universal indicator to a test-tube containing water?

Q22. Which chemical is injected into the skin of a person during an ant's sting?

Q23. Which chemical is injected into the skin of a person during the nettle leaf hair sting?

Q24. How can the effect of these stings be neutralised?

PRACTICE EXERCISE 2.6

Q1. Explain the pH change as the cause of tooth decay. How can tooth decay caused by pH change be prevented?

Q2. Explain how pH change in the lake water can endanger the lives of aquatic animals (like fish). What can be done to lessen the danger to the lives of aquatic animals in the lake?

Q3. What happens during a bee sting? What is its remedy?

Q4. What happens during a wasp sting? What is its remedy?

Q5. Why is it wrong to treat a bee sting with vinegar?

Q6. Why is it wrong to treat a wasp sting with baking soda solution?

Q7. What does the pH of a solution signify? Three solutions A, B and C have pH values of 6, 4 and 10 respectively. Which of the solutions is highly acidic?

Q8. A farmer has found that the pH of soil in his fields is 4.2. Name any two chemical materials which he can mix with the soil to adjust its pH.

Q9. The pH values of six solutions A to F are given below:
 $A = 0, B = 11, C = 6, D = 3,$
 $E = 13, F = 8$
 Name the acids or alkalis used to make (i) car batteries (ii) explosives (iii) soaps (iv) fertilisers.

Q10. The pH of a cold drink is 5. What will be its action on blue and red litmus solutions?

Q11. The pH values of three acids A, B and C having equal molar concentrations are 5.0, 2.8 and 3.5 respectively. Arrange these acids in order of the increasing acid strengths.

Q12. Under what soil conditions do you think a farmer would treat the soil of his fields with quicklime (calcium oxide), or slaked lime (calcium hydroxide) or chalk (calcium carbonate)?

Q13. Which acid is produced in our stomach? What happens if there is an excess of acid in the stomach? How can its effect be cured?

Q14. The soil in a field is highly acidic. Name two materials which can be added to this soil to reduce its acidity. Give the reason for your choice.

Q15. What is meant by strong bases and weak bases? Classify the following into strong bases and weak bases:
 $\text{NH}_4\text{OH}, \text{Ca}(\text{OH})_2, \text{NaOH}, \text{ZnOH}$
 $\text{KOH}, \text{Mg}(\text{OH})_2, \text{Al}(\text{OH})^2, \text{Fe}(\text{OH})_2$

Q16. What ions are present in the solutions of following substances? (write the symbols only)
 (i) Hydrochloric acid
 (ii) Nitric acid (iii) Sulphuric acid
 (iv) Sodium hydroxide
 (v) Potassium hydroxide
 (vi) Magnesium hydroxide

Q17. What would you expect the pH of pure water to be?

Q18. What colour would the universal indicator show in an aqueous solution of sugar? Why?

Q19. A sample of rain water turned universal indicator paper yellow. What would you expect its pH to be? Is it a strong or a weak acid?

Q20. What do you think will be the pH in the stomach of a person suffering from indigestion: less than 7 or more than 7?

Q21. What do you think will be the pH of an antacid solution: less than 7 or more than 7?

Q22. How does an antacid work?

Q23. Name two common antacids.

PRACTICE EXERCISE 2.7

Q1. Separate the following into substances having pH values above and below 7. How do these influence litmus paper?

(a) Lemon juice (b) Stomach juices
(c) Solution of washing soda
(d) Toothpaste (e) Vinegar

Q2. Do basic solutions also have H^+ (aq) ions? If yes, then why are they basic?

Q3. When a solution becomes more acidic, does the pH get higher or lower?

Q4. Define an acid and a base. Give two examples of each.

Q5. Give the names and formulae of two strong bases and two weak bases.

Q6. What type of ions are formed?

(i) when an acid is dissolved in water?
(ii) when a base (or alkali) is dissolved in water?

Q7. Write the neutralisation reaction between acids and bases in terms of the ions involved.

Q8. Write any two important uses of bases.

Q9. What happens when zinc granules are heated with sodium hydroxide solution? Write equation of the reaction which takes place.

Q10. What happens when bases react with non metal oxides? Explain with the help of an example. What does this reaction tell us about the nature of non-metal oxides?

Q11. What effect does the concentration of H^+ (aq) ions have on the nature of a solution?

Q12. What effect does the concentration of OH^- ions have on the nature of a solution?

Q13. Someone put some universal indicator paper into vinegar. The pH is 3. What does this tell you about the vinegar?

Q14. Someone put some universal indicator paper onto wet soap. The pH is 8. What does this tell you about the soap?

Q15. State whether a solution is acidic, alkaline or neutral if its pH is:

(i) 9 (ii) 4
(iii) 7 (iv) 1
(v) 10 (vi) 3

Q16. A milkman adds a very small amount of baking soda to fresh milk.

(a) Why does he shift the pH of the fresh milk from 6 to slightly alkaline?
(b) Why does this milk take a long time to set as curd?

Q17. Which of the following elements would form oxides which would indicate pH values less than seven, using moist pH paper?

Magnesium, Carbon, Sulphur, Hydrogen, Copper

Q18. The pH values of six solutions A, B, C, D E and F are given below:

A	1	B	5
C	7	D	11
E	13	F	0

Which solution is (i) weakly alkaline (ii) neutral (iii) strongly acidic (iv) strongly alkaline, and (v) weakly acidic?

Q19. Potatoes grow well on Anhad's farm which has soil with a pH of 5.5. Anhad decides to add lot of lime to soil so that he can grow broccoli in the same farm:

(a) Do potatoes grow better in acidic or alkaline soil?
(b) Does broccoli grow better in acidic or alkaline soil?

Q20. Here are some results of solutions tested with universal indicator paper:

Sulphuric acid	Red
Metal polish	Dark blue
Washing-up liquid	Yellow
Milk of magnesia	Light blue
Oven cleaner	Purple
Car battery acid	Pink

Arrange the solutions in order of their increasing pH values (starting with the one with the lowest pH)

PRACTICE EXERCISE 2.8

Q1. Solution A turns universal indicator blue to purple whereas solution B turns universal indicator orange to red.

- What will be the action of solution A on litmus?
- What will be action of solution B on litmus?
- Name any two substances which can give solutions like A.
- Name any two substances which can give solutions like B.
- What sort of reaction takes place when solution A reacts with solution B?

Q2. A first-aid manual suggests that vinegar should be used to treat wasp stings and baking soda for bee sting. What does this information tell you about the chemical nature of:

- wasp stings? (b) bee stings?

Q3. (a) Explain why the pH in a person's mouth becomes lower after each meal.
 (b) What damage could be caused while the pH is low?
 (c) How could the person change his eating habits to lessen changes of suffering from tooth decay?

Q4. A group of students measured the pH of some substances they found in their homes. Their results are given in the following table:

Name of Substance	pH value of substance
Apples	3.0
Baking soda	8.5
Black coffee	5.8
Household ammonia	12.0
Lemon juice	2.5
Milk	6.5
Salt	7.0
Sugar	7.0
Toothpaste	9.0
Vinegar	3.0
Washing soda	11.5

- What would the students have used to measure the pH?
- Which solution is the most acidic?
- Which solution is the most alkaline?
- Which solutions are neutral?
- Which solution can be used to treat wasp stings?
- Which solution can be used to treat bee stings?

Q5. Hydrochloric acid reacts with a metal X to form a gas Y which burns with a 'pop' sound. Sodium hydroxide solution also reacts with the same metal X (on heating) to form the same gas Y.

- Name X and Y.
- Write the chemical equation of the reaction of metal X with (i) hydrochloric acid, and (ii) sodium hydroxide solution.

Q6. (a) An aqueous solution has a pH value of 7.0. Is this solution acidic, basic or neutral.
 (b) If H^+ concentration of the solution is $1 \times 10^{-2} \text{ mol L}^{-1}$, what is the pH value for it?
 (c) Which has a higher pH value, 1 M HCl or 1 M NaOH solution.

Q7. Name the raw materials that are required for the manufacture of washing soda by solvay process. Describe the chemical reactions involved in the process.

Q8. A small amount of hydrochloric acid is always produced in the stomach. Is it useful or harmful and in what way? If excess of acid is produced in the stomach, what do we do.

Q9. Give examples of three acids which are useful in our eatables stating the use of which they are put.

PRACTICE EXERCISE 2.9

Q1. One of the following is a medicine for indigestion. This is:

- (a) sodium hydroxide
- (b) manganese hydroxide
- (c) magnesium hydroxide
- (d) potassium hydroxide

Q2. Bee sting contains:

- (a) an acidic liquid
- (b) a salt solution
- (c) an alkaline liquid
- (d) an alcohol

Q3. Wasp sting contains:

- (a) a sugar solution
- (b) an acidic liquid
- (c) a salt solution
- (d) an alkaline liquid

Q4. One of the following does not inject an acidic liquid into the skin through its sting. This is:

- (a) honey bee
- (b) ant
- (c) wasp
- (d) nettle leaf hair

Q5. A solution turns red litmus blue. Its pH is likely to be:

- (a) 1
- (b) 4
- (c) 5
- (d) 10

Q6. A solution turns blue litmus red. Its pH is likely to be:

- (a) 7
- (b) 5
- (c) 8
- (d) 14

Q7. A solution turns phenolphthalein indicator pink. The most likely pH of this solution will be:

- (a) 6
- (b) 4
- (c) 9
- (d) 7

Q8. The colour of methyl orange indicator in a solution is yellow. The pH of this solution is likely to be:

- (a) 7
- (b) less than 7
- (c) 0
- (d) more than 7

Q9. Bee stings can be treated with:

- (a) vinegar
- (b) sodium hydrogencarbonate
- (c) potassium hydroxide
- (d) lemon juice

Q10. Wasp stings can be treated with:

- (a) baking soda
- (b) vinegar

- (c) washing soda

- (d) milk of magnesia

Q11. It has been found that rubbing vinegar on the stung area of the skin of a person gives him relief. The person has been stung by:

- (a) wasp
- (b) ant
- (c) honey bee
- (d) nettle leaf hair

Q12. Fresh milk has a pH of 6. When milk changes into curd, the pH value will:

- (a) become 7
- (b) become less than 6
- (c) become more than 7
- (d) remain unchanged

Q13. The acid produced naturally in our stomach is:

- (a) acetic acid
- (b) citric acid
- (c) hydrochloric acid
- (d) sulphuric acid

Q14. The daffodil plants grow best in a soil having a pH range of 6.0 to 6.5. If the soil in a garden has a pH of 4.5, which substance needs to be added to the soil in order to grow daffodils?

- (a) salt
- (b) lime
- (c) sand
- (d) compost

Q15. Nettle leaf strings can be treated with?

- (a) NaOH
- (b) Fe(OH)₂
- (c) dock plant
- (d) HCl

Q16. Which of the following compounds give large hydroxide ions in aqueous solution:

- (a) Ca(OH)₂
- (b) NH₄OH
- (c) Mg(OH)₂
- (d) NaOH

Q17. Which of the following compounds are good electrolytes:

- (a) NaOH
- (b) NH₄OH
- (c) Ca(OH)₂
- (d) ZnOH

Q18. Which of the following compounds in aqueous solution is the best conductor of electricity:

- (a) Mg(OH)₂
- (b) ZnOH
- (c) Ca(OH)₂
- (d) KOH

PRACTICE EXERCISE 2.10

Q1. What is the chemical formula of (a) baking soda, and (b) washing soda?

Q2. Write the chemical formula of (i) soda ash, and (ii) sodium carbonate decahydrate.

Q3. State whether the following statement is true or false:
Copper sulphate crystals are always wet due to the presence of water of crystallisation in them.

Q4. Which of the following salt has a blue colour and why? $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ or CuSO_4

Q5. What would be the colour of litmus in a solution of sodium carbonate?

Q6. State the common and chemical names of the compound formed when plaster of Paris is mixed with water.

Q7. With which substance should chlorine be treated to get bleaching powder?

Q8. What is the commercial name of calcium sulphate hemihydrate?

Q9. Name the product formed when Cl_2 and H_2 produced during the electrolysis of brine are made to combine.

Q10. Name a calcium compound which hardens on wetting with water.

Q11. Name a sodium compound which is a constituent of many dry soap powders.

Q12. Name a metal carbonate which is soluble in water.

Q13. Name an acid which is present in baking powder.

Q14. Name the metal whose carbonate is known as washing soda.

Q15. Which compound is used as an antacid in medicine: NaHCO_3 or Na_2CO_3 ?

Q16. What is the common name of (a) NaHCO_3 , and (b) $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$?

Q17. Write the chemical name and formula of (a) common salt, and (b) caustic soda.

Q18. What are the two main ways in which common salt (sodium chloride) occurs in nature?

Q19. What are the two main ways in which common salt (sodium chloride) occurs in nature?

Q20. How is common salt obtained from sea-water?

Q21. Why is sodium chloride required in our body?

Q22. Name three chemicals made from common salt (or sodium chloride).

Q23. Give any two uses of common salt (sodium chloride).

Q24. What name is given to the common salt which is mined from underground deposits? How was this salt formed?

Q25. Name the salt which is used as a preservative in pickles, and in curing meat and fish.

Q26. Name the raw material used for the production of caustic soda.

Q27. The electrolysis of an aqueous solution of sodium chloride gives us three products. Name them.

Q28. During the electrolysis of a saturated solution of sodium chloride, where is:
(a) chlorine formed?
(b) hydrogen formed?
(c) sodium hydroxide formed?

Q29. Fill in the following blanks:
(a) Common salt is obtained from sea-water by the process of
(b) Rock salt is mined just like
(c) Chemical formula of washing soda is
(d) Sodium hydrogencarbonate is soda whereas sodium carbonate is soda.
(e) The chemical formula of plaster of Paris is

Q30. Complete and balance the following chemical equations:
(a) $\text{NaCl} \text{ (aq)} + \text{H}_2\text{O} \text{ (l)} \xrightarrow[\text{H}_2\text{O}]{\text{E, heat}}$
(b) $\text{NaHCO}_3 \xrightarrow{\text{H}_2\text{O}}$

PRACTICE EXERCISE 2.11

Q1. What is washing soda? State two properties and two uses of washing soda.

Q2. Write the formulae of sodium chloride and sodium chloride and sodium carbonate. Explain why an aqueous solution of sodium chloride is neutral but an aqueous solution of sodium carbonate is basic (or alkaline). Write chemical equations of the reactions involved.

Q3. Write the chemical formula of ammonium chloride. Explain why an aqueous solution of ammonium chloride is acidic in nature? Illustrate your answer with the help of a chemical equation.

Q4. What is baking soda? Write the chemical name of baking soda. Give the important uses of baking soda. How does baking soda differ chemically from washing soda?

Q5. Describe how sodium hydrogencarbonate (baking soda) is produced on a large scale. Write equation of the reaction involved.

Q6. What happens when a cold and concentrated solution of sodium chloride reacts with ammonia and carbon dioxide? Write the chemical equation of the reaction which takes place.

Q7. What is meant by “water of crystallisation” in a substance? Explain with an example.

Q8. How would you show that blue copper sulphate crystals contain water of crystallisation?

Q9. Explain how anhydrous copper sulphate can be used to detect the presence of moisture (water) in a liquid.

Q10. What is the common name of sodium hydrogencarbonate?

Q11. What happens when a solution of sodium hydrogencarbonate is heated? Write equation of the reaction involved.

Q12. Explain why, sodium hydrogencarbonate is used as an antacid.

Q13. What will happen if heating is not controlled while preparing plaster of Paris?

Q14. Write an equation to show the reaction between plaster of Paris and water.

Q15. What happens when copper sulphate crystals are heated strongly? Explain with the help of an equation.

Q16. What happens when a few drops of water are added to anhydrous copper sulphate? Explain with the help of an equation.

Q17. Name two constituents of baking powder.

Q18. How does baking powder differ from baking soda?

Q19. Explain the action of baking powder in the making of cake (or bread). Write equation of the reaction involved.

Q20. What is the chemical name of bleaching powder?

Q21. What is the chemical formula of bleaching powder?

Q22. What are the materials used for the preparation of bleaching powder?

Q23. State one use of bleaching powder (other than bleaching).

Q24. What does a soda-acid type fire extinguisher contain? How does it work? Explain the working of a soda-acid fire extinguisher with the help of a labelled diagram.

Q25. Name a sodium compound used for softening hard water.

Q26. Which compound of calcium is used for disinfecting drinking water supply?

Q27. Name a metal compound which has detergent properties (cleansing properties).

Q28. Name one compound of calcium which is used for removing the colour of a coloured cloth.

Q29. State a peculiar (or remarkable) property of plaster of Paris.

Q30. Name the substance obtained by the action of chlorine on solid (dry) slaked lime.

PRACTICE EXERCISE 2.12

Q1. What is gypsum? What happens when gypsum is heated to 100°C (373 K)?

Q2. Name a sodium compound which is used for making borax and glass.

Q3. Name the compound which is used in hospitals for setting fractured bones.

Q4. Which is the real bleaching agent present in bleaching powder?

Q5. What is “baking powder”? How does it make the cake soft and spongy?

Q6. In addition to sodium hydrogencarbonate, baking powder contain a substance X. Name the substance X. What is the role of substance X in the baking powder?

Q7. State two uses each of the following compounds:
(a) Sodium hydroxide (b) Chlorine
(c) Hydrogen (b) Hydrochloric acid

Q8. What is the common name of the compound $CaOCl_2$?

Q9. Name the raw material used for the preparation of plaster of Paris.

Q10. Which property of plaster of Paris is utilised in making casts for broken limbs in hospitals?

Q11. Explain why chlorine is used for sterilising drinking water supply.

Q12. What happens when a concentrated solution of sodium chloride (brine) is electrolysed? Write the equation of the reaction involved.

Q13. Why is the electrolysis of a concentrated solution of sodium chloride known as chlor alkali process?

Q14. Name three products of the chlor-alkali process. State two uses of each of these products.

Q15. Describe how washing soda is produced starting from sodium chloride (common salt). Write equations of all the reactions involved.

Q16. State whether an aqueous solution of washing soda is acidic or alkaline? Give reason for your answer.

Q17. What is meant by saying that washing soda has detergent properties?

Q18. Give two important uses of washing soda (or sodium carbonate).

Q19. What is bleaching powder? How is bleaching powder prepared? Write chemical equation of the reaction involved in the preparation of bleaching powder.

Q20. What happens when bleaching powder reacts with dilute sulphuric acid? Give equation of the reaction involved.

Q21. State two important uses of bleaching powder.

Q22. What is plaster of Paris? Write the chemical formula of plaster of Paris.

Q23. How is plaster of Paris prepared? Write chemical equation of the reaction involved.

Q24. Explain why plaster of Paris should be stored in a moisture-proof container.

Q25. State two important uses of plaster of Paris.

Q26. What is a salt? Give the names and formulae of any two salts. Also name the acids and bases from which these salts may be obtained.

Q27. What is meant by ‘a family of salts’? Explain with examples.

Q28. What is meant by ‘hydrated’ and ‘anhydrous’ salts? Explain with examples.

Q29. Write the names, formulae and colour of any two hydrated salts.

Q30. What will be the colour of litmus in an aqueous solution of ammonium chloride salt?

Q31. P and Q are aqueous solutions of sodium chloride and sodium hydroxide, respectively. Which of these will turn:
(a) blue litmus red?
(b) red litmus blue?

Q32. The metal salt A is blue in colour. When salt A is heated strongly over a burner, then a substance B is eliminated and a white powder C is left behind. When a few drops of a liquid D are added to powder C, it becomes blue again. What could be A, B and C?

PRACTICE EXERCISE 2.13

Q1. When the concentrated aqueous solution of substance X is electrolysed, then NaOH , Cl_2 and H_2 are produced. Name the substance X. What is the special name of this process?

Q2. Consider the following substances:
 NaCl , $\text{Ca}(\text{OH})_2$, NaHCO_3 , NH_3 ,
 Na_2CO_3 , H_2O , Cl_2 , CO_2 , CaSO_4 ,
 $2\text{H}_2\text{O}$, $2\text{CaSO}_4 \cdot \text{H}_2\text{O}$, CaOCl_2

- Which two substances combine to form bleaching powder?
- Which four substances are utilised in the production of washing soda?
- Which compound represents plaster of Paris?
- Which compound is a part of baking powder?
- Which compound is used as an antacid?

Q3. Give one example each of a salt which gives an aqueous solution having:

- pH less than 7
- pH equal to 7
- pH more than 7

Q4. A compound X which is prepared from gypsum has the property of hardening when mixed with a proper quantity of water.

- Identify the compound X
- Write the chemical equation for its preparation.
- For what purpose is it used in hospitals?

Q5. Consider the following salts:
 Na_2CO_3 , NaCl , NH_4Cl ,
 CH_3COONa , K_2SO_4 , $(\text{NH}_4)_2\text{SO}_4$

- acidic solutions?
- neutral solutions?
- basic solutions (or alkaline solutions)?

Q6. A white powdery substance having strong smell of chlorine is used for disinfecting drinking water supply at waterworks. Identify the substance. Give its chemical name and write the chemical reaction for its preparation.

Q7. A salt X when dissolved in distilled water gives a clear solution which turns red litmus blue. Explain the phenomenon.

Q8. A person found that the cake prepared by him is hard and small in size. Which ingredient has he forgotten to add that would have caused the cake to rise and become light? Explain your answer.

Q9. A white chemical compound becomes hard on mixing with proper quantity of water. It is also used in surgery to maintain joints in a fixed position. Name the chemical compound.

Q10. When chlorine and sodium hydroxide being produced during the electrolysis of brine are allowed to mix, a new chemical is formed. Name this chemical and write its uses.

Q11. Write the name and formula of one salt each which contains:

- two molecules of water of crystallisation
- five molecules of water of crystallisation
- ten molecules of water of crystallisation

Q12. How many molecules of water of crystallisation (per formula unit) are present in:

- copper sulphate crystals?
- washing soda? (c) gypsum?

Q13. pH of NH_4Cl and CuSO_4 solution will be:

- 7
- > 7
- < 7
- 0

Q14. On electrolysis of brine solution, the product formed:

- Na , Cl_2
- H_2 , Cl_2 , O_2
- NaOH , H_2Cl_2
- NaOH , Cl_2 , O_2

Q15. Aqueous solution of Na_2CO_3 is

- acidic
- basic
- neutral
- amphoteric

Q16. Plaster of paris on mixing with water set to form:

- $\text{CaSO}_4 \cdot \text{H}_2\text{O}$
- $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$
- $\text{CaSO}_4 \cdot 0.5\text{H}_2\text{O}$
- $\text{CaSO}_4 \cdot 5\text{H}_2\text{O}$

PRACTICE EXERCISE 2.14

Q1. The salt which will give an acidic solution on dissolving in water is:

- KCl
- NH_4Cl
- Na_2CO_3
- CH_3COONa

Q2. One of the following salts will give an alkaline solution on dissolving in water. This is:

- Na_2CO_3
- Na_2SO_4
- NaCl
- $(\text{NH}_4)_2\text{SO}_4$

Q3. The salt which will give a neutral solution on dissolving in water will be:

- CH_3COONa
- NH_4Cl
- KCl
- Na_2CO_3

Q4. The products of chlor-alkali process are:

- NaCl, Cl_2 and H_2
- H_2 , Cl_2 and NaOH
- Cl_2 , Na_2CO_3 and H_2O
- NaOH, Cl_2 and HCl

Q5. The number of molecules of water of crystallisation present in washing soda crystals is:

- five
- two
- ten
- seven

Q6. The salt whose aqueous solution will turn blue litmus to red is:

- ammonium sulphate
- sodium acetate
- sodium chloride
- potassium carbonate

Q7. The aqueous solution of one of the following salts will turn red litmus to blue. This salt is:

- potassium sulphate
- sodium sulphate
- sodium chloride
- potassium carbonate

Q8. The salt whose aqueous solution will have no effect on either red litmus or blue litmus is

- potassium sulphate
- sodium carbonate
- ammonium sulphate
- sodium acetate

Q9. The aqueous solution of one of the following salts will turn phenolphthalein indicator pink. This salt is:

- KCl
- K_2SO_4
- K_2CO_3
- KNO_3

Q10. The formula of baking soda is:

- K_2CO_3
- KHCO_3
- NaHCO_3
- Na_2CO_3

Q11. Which of the following is treated with chlorine to obtain bleaching powder?

- CaSO_4
- $\text{Ca}(\text{OH})_2$
- $\text{Mg}(\text{OH})_2$
- KOH

Q12. Plaster of Paris is prepared by heating one of the following to a temperature of 100°C . This is:

- $\text{CaSO}_3 \cdot 2\text{H}_2\text{O}$
- $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$
- $\text{CaCO}_3 \cdot 2\text{H}_2\text{O}$
- $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$

Q13. A salt whose aqueous solution will have a pH of more than 7 will be:

- K_2CO_3
- K_2SO_4
- NaCl
- NH_4Cl

Q14. A salt is dissolved in water and the pH of this solution is measured with a universal indicator paper. If the pH of solution is less than 7, the salt is most likely to be:

- CH_3COONa
- Na_2CO_3
- KCl
- NH_4Cl

Q15. Which of the following salts will give an aqueous solution having pH of almost 7?

- NH_4NO_3
- NH_4Cl
- CaCl_2
- KCl

Q16. The common name of CaOCl_2 is:

- washing soda
- bleaching powder
- caustic soda
- potash soda

Q16. The chemical formula of caustic soda is:

- CaCl_2
- NH_4OH
- NaOH
- ZnCO_3

Q17. The number of molecules of water of crystallization present in gypsum:

- 3
- 4
- 2
- 5